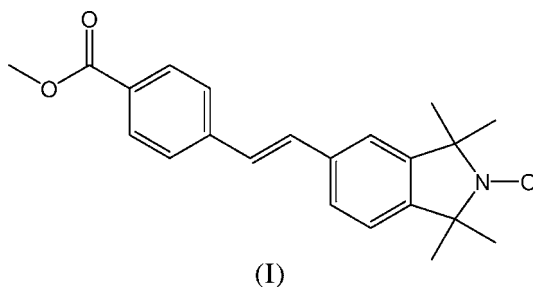


5-[(*E*)-2-(4-Methoxycarbonylphenyl)ethenyl]-
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Key indicators

Single-crystal X-ray study
T = 150 K
Mean $\sigma(\text{C}-\text{C})$ = 0.002 Å
R factor = 0.042
wR factor = 0.114
Data-to-parameter ratio = 19.2For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.The asymmetric unit of the title compound, C₂₂H₂₄NO₃,
comprises two molecules with an *E* stereochemical configura-
tion at the ethylene groups.Received 14 July 2006
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Comment

We are currently investigating the chemistry of paramagnetic
profluorescent organic compounds. Of particular interest are
the isoindoline class of nitroxide radicals (aminoxyls), such as
TMIO (Smith *et al.*, 2000, 2002), as these display superior
electron paramagnetic resonance (EPR) linewidths which
results in increased accuracy in EPR oximetry (Bolton *et al.*,
1993; Gillies *et al.*, 1994; Bottle *et al.*, 1999, 2000; Shen *et al.*,
2002). Furthermore, the isoindoline system offers the oppor-
tunity to synthesize more complex structures through substi-
tution of the aromatic ring, thus leading to the potential for
advanced applications as radical sensors (Micallef *et al.*, 2005).
The title compound, (I), was prepared as part of these ongoing
investigations (Keddie *et al.*, 2005).The asymmetric unit of (I), illustrated in Fig. 1, contains two
molecules, both of which have an *E* (*trans*) stereochemical
configuration at the ethylene bridge. The synthesis of this
compound was performed by palladium-catalyzed Heck
coupling of bromo-TMIO and methyl-4-vinylbenzoate. Both
of these components are relatively bulky and were therefore
expected to favour formation of the *trans* alkene product
(Braese & Meijere, 2002). The nitroxide N—O bond lengths
(Table 1) in the two molecules are typical for isoindoline
nitroxide radicals (Busfield *et al.*, 1986; Fairhurst *et al.* 1996).In the crystal structure of (I), the molecules are arranged in
a zigzag-like pattern that propagates along the *c* axis. Inter-
estingly, despite the obvious potential for intermolecular π – π
interactions to occur between the benzene rings of the mol-
ecules of (I), there are no such interactions of significance
apparent in this structure. The only directional interactions
identified are methyl-C—H··· π contacts (C21—H21C···C3ⁱ–
C8_{plane}ⁱ = 2.866 Å, C22—H22C···C33–C38_{plane} = 2.688 Å,
C43—H43A···C25ⁱ–C30_{plane}ⁱ = 3.070 Å and C44—H44C···
C11ⁱⁱ–C16_{plane}ⁱⁱ = 2.852 Å).

Experimental

The title compound was prepared by Heck coupling of 5-bromo-TMIO (Micallef *et al.*, 1999) and methyl-4-vinylbenzoate, according to the literature procedure reported by us (Keddie *et al.*, 2005). Slow evaporation of a 70:30 hexane–ethyl acetate solution of (I) afforded the large orange prismatic crystals used for this study.

Crystal data

$C_{22}H_{24}NO_3$	$Z = 8$
$M_r = 350.42$	$D_x = 1.222 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/c$	Mo $K\alpha$ radiation
$a = 11.6300 (19) \text{ \AA}$	$\mu = 0.08 \text{ mm}^{-1}$
$b = 13.432 (2) \text{ \AA}$	$T = 150 (2) \text{ K}$
$c = 24.407 (4) \text{ \AA}$	Prism, orange
$\beta = 91.976 (3)^\circ$	$0.41 \times 0.37 \times 0.19 \text{ mm}$
$V = 3810.5 (11) \text{ \AA}^3$	

Data collection

Bruker SMART 1000 CCD area-detector diffractometer	37652 measured reflections
ω scans	9218 independent reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 1999)	7037 reflections with $I > 2\sigma(I)$
$T_{\min} = 0.91, T_{\max} = 0.99$	$R_{\text{int}} = 0.030$
	$\theta_{\max} = 28.4^\circ$

Refinement

Refinement on F^2	$w = 1/[\sigma^2(F_o^2) + (0.0565P)^2 + 0.9099P]$
$R[F^2 > 2\sigma(F^2)] = 0.042$	where $P = (F_o^2 + 2F_c^2)/3$
$wR(F^2) = 0.114$	$(\Delta/\sigma)_{\max} < 0.001$
$S = 1.01$	$\Delta\rho_{\max} = 0.27 \text{ e \AA}^{-3}$
9218 reflections	$\Delta\rho_{\min} = -0.23 \text{ e \AA}^{-3}$
479 parameters	
H-atom parameters constrained	

Table 1

Selected bond lengths (\AA).

C9–C10	1.3328 (17)	N1–O3	1.2758 (13)
C31–C32	1.3297 (18)	N2–O6	1.2728 (14)

C-bound H atoms were included in idealized positions and refined using a riding model. Methine, aromatic and methyl C–H bond lengths were fixed at 0.95, 0.95 and 0.98 \AA , respectively. $U_{\text{iso}}(\text{H})$ values were fixed at $1.2U_{\text{eq}}(\text{C})$ for methine and aromatic H atoms, and at $1.5U_{\text{eq}}(\text{C})$ for methyl H atoms.

Data collection: SMART (Siemens, 1995); cell refinement: SAINT (Siemens, 1995); data reduction: SAINT and XPREP (Siemens, 1995); program(s) used to solve structure: SIR97 (Altomare *et al.*, 1999); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: ORTEP-3 (Farrugia, 1997) and WinGX32 (Farrugia, 1999); software used to prepare material for publication: enCIFer (Allen *et al.*, 2004).

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References

Allen, F. H., Johnson, O., Shields, G. P., Smith, B. R. & Towler, M. (2004). *J. Appl. Cryst.* **37**, 335–338.

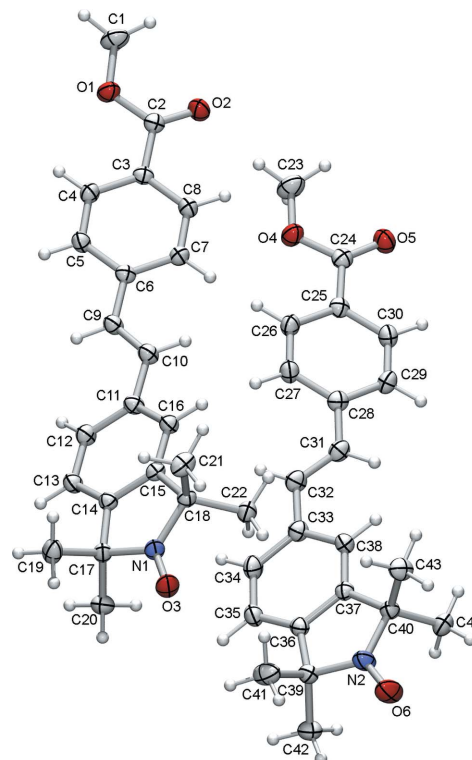


Figure 1

The two crystallographically independent molecules of (I), with the atom-numbering scheme. Displacement ellipsoids are drawn at the 50% probability level.

- Altomare, A., Burla, M. C., Camalli, M., Casciarano, G. L., Giacovazzo, C., Guagliardi, A., Moliterni, A. G. G., Polidori, G. & Spagna, R. (1999). *J. Appl. Cryst.* **32**, 115–119.
- Bolton, R., Gillies, D. G., Sutcliffe, L. H. & Wu, X. (1993). *J. Chem. Soc. Perkin Trans. 2*, pp. 2049–2052.
- Bottle, S. E., Gillies, D. G., Hughes, D. L., Micallef, A. S., Smirnov, A. I. & Sutcliffe, L. H. (2000). *J. Chem. Soc. Perkin Trans. 2*, pp. 1285–1291.
- Bottle, S. E., Gillies, D. G., Micallef, A. S., Reid, D. A. & Sutcliffe, L. H. (1999). *Magn. Reson. Chem.* **37**, 730–734.
- Braese, S. & Meijere, A. (2002). *Handbook of Organopalladium Chemistry for Organic Synthesis*, edited by E.-I. Negishi, Vol. 1, pp. 1123–1132. Chichester: John Wiley and Sons.
- Busfield, W. K., Engelhardt, L. M., Healy, P. C., Jenkins, I. D., Thang, S. H. & White, A. H. (1986). *Aust. J. Chem.* **39**, 357–365.
- Fairhurst, S. A., Gillies, D. G., Smith, G. W., Sutcliffe, L. H. & Wu, X. (1996). *J. Mol. Struct.* **375**, 105–115.
- Farrugia, L. J. (1997). *J. Appl. Cryst.* **30**, 565.
- Farrugia, L. J. (1999). *J. Appl. Cryst.* **32**, 837–838.
- Gillies, D. G., Sutcliffe, L. H. & Wu, X. (1994). *J. Chem. Soc. Faraday Trans. 90*, 2345–2349.
- Keddie, D. J., Johnson, T. E., Arnold, D. P. & Bottle, S. E. (2005). *Org. Biomol. Chem.* **3**, 2593–2598.
- Micallef, A. S., Blinco, J. P., George, G. A., Reid, D. A., Rizzardo, E., Thang, S. H. & Bottle, S. E. (2005). *Polym. Degrad. Stab.* **89**, 427–435.
- Micallef, A. S., Bott, R. C., Bottle, S. E., Smith, G., White, J. M., Matsuda, K. & Iwamura, H. (1999). *J. Chem. Soc. Perkin Trans. 2*, pp. 65–71.
- Sheldrick, G. M. (1997). *SHELXL97*. University of Göttingen, Germany.
- Sheldrick, G. M. (1999). *SADABS*. University of Göttingen, Germany.
- Shen, J., Bottle, S. E., Khan, N., Grinberg, O., Reid, D. A., Micallef, A. S. & Swartz, H. (2002). *Appl. Magn. Reson.* **22**, 357–368.
- Siemens (1995). *SMART, SAINT and XPREP*. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
- Smith, C. D., Bartley, J. P., Bottle, S. E., Micallef, A. S. & Reid, D. A. (2000). *J. Mass Spectrom.* **35**, 607–611.
- Smith, C. D., Bott, R. C., Bottle, S. E., Micallef, A. S. & Smith, G. (2002). *J. Chem. Soc. Perkin Trans. 2*, pp. 533–537.